

specific heat chem worksheet pdf

5. Copper has a specific heat of $0.385 \text{ J/(g}\cdot\text{)}^\circ\text{C}$. A piece of copper absorbs 5000 J of energy and undergoes a temperature change from 100°C to 200°C . What is the mass of the piece of copper? $q = 5000 \text{ J}$ $m = ?$ $c = 0.385 \text{ J/g}\cdot\text{)}^\circ\text{C}$ $\Delta T = 200^\circ\text{C} - 100^\circ\text{C} = 100^\circ\text{C}$ $m = 129.87 \text{ g}$ Endothermic or exothermic? Endothermic 6.

Worksheet- Calculations involving Specific Heat

b) the lowest heat capacity? 6. Here are the heat capacities of the four substances: $4.18 \text{ J/g}\cdot\text{)}^\circ\text{C}$, $1.00 \text{ J/g}\cdot\text{)}^\circ\text{C}$, $0.80 \text{ J/g}\cdot\text{)}^\circ\text{C}$, & $0.60 \text{ J/g}\cdot\text{)}^\circ\text{C}$. Match & then label each substance with its specific heat capacity on the graph. 7. If something has a high specific heat capacity will it take a lot of heat or a little heat to change its temperature? Explain.

Name: Per: Worksheet- Introduction to Specific Heat Capacities

4) What are the units of specific heat? 5) What is the symbol for specific heat? 6) Look at the specific heat capacity chart on the first page of this worksheet.

Specific Heat Calculations Worksheet Name: Chemistry

A silver ring has a mass of 138.45 g. How many calories of heat are required to increase the temperature from 11.8°C to 162.5°C ? 6. A heat energy of 645 J is applied to a sample of glass with a mass of 28.4 g. Its temperature increases from 11.6°C to 15.5°C . Calculate the specific heat of glass. 7.

Specific Heat - California State University, Northridge

Chemistry*Temperature&SpecificHeat*Worksheet*Answer Key TemperatureConversions! 1. Complete!the!table!below:!!!! ! 2" 3" 4"

Chemistry*Temperature&SpecificHeat*Worksheet* Answer Key

heat of the solution is $4.18 \text{ J/g}\cdot\text{)}^\circ\text{C}$, that its density is 1.00 /mL , and that the calorimeter itself absorbs a negligible amount of heat, calculate the amount of heat absorbed for the reaction.

Specific Heat and Heat Capacity Worksheet

Chemistry II Enthalpy Worksheet Name _____ As we have studied heat capacity and specific heat capacity, we've learned how heat can be exchanged between objects. Heat can also be exchanged during a chemical reaction. In an exothermic chemical reaction, heat is a product; while in an endothermic chemical reaction, heat is a reactant.

Chemistry II Enthalpy Worksheet Name

17. What is the specific heat of a substance if removing 95.0 J of energy decreases the temperature of 5.0 g of the substance from 23.5°C to 18.1°C ? Calorimetry Calculations 18. A chemical reaction takes place in 255 g of water. During the reaction the temperature increased from 22.8°C to 31.2°C .

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