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4 Chapter 1 Introduction to Chemistry Section 11.1.1 Figure 1.1 Everything in the universe, including particles in space and things around you, is composed of matter. Objectives Define substance. Explain the formation and

Chapter 1: Introduction to Chemistry

As you begin your study of chemistry—the study of matter and the changes that it undergoes—you are probably asking yourself, “Why is chemistry important to me?” The answer to this question can be illustrated by real-life events that involve two discoveries.

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CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. homogeneous a. sugar homogeneous d. plastic wrap homogeneous b. iron filings heterogeneous e. cement sidewalk heterogeneous c. granola bar 2.

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â€¢ Matter is composed of extremely small particles called atoms. â€¢ Atoms are indivisible and indestructible. â€¢ Atoms of a given element are identical in size, mass, and chemical properties.

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Chapter 21: Hydrocarbons

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A change of matter from one form to another without a change iâ€! A change in which one or more substances combine or break aparâ€! The process by which molecules at the surface of a liquid absoâ€!

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284 Chapter 9 â€¢ Chemical Reactions ... conservation of mass states that in a chemical change, matter is neither created nor destroyed. Chemical equations must show that matter is conserved during a reaction. Skeleton equations lack that information. Look at Figure 9.4.

Chapter 9: Chemical Reactions

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